



Bi-Atom Electrocatalyst for Electrochemical Nitrogen Reduction Reactions

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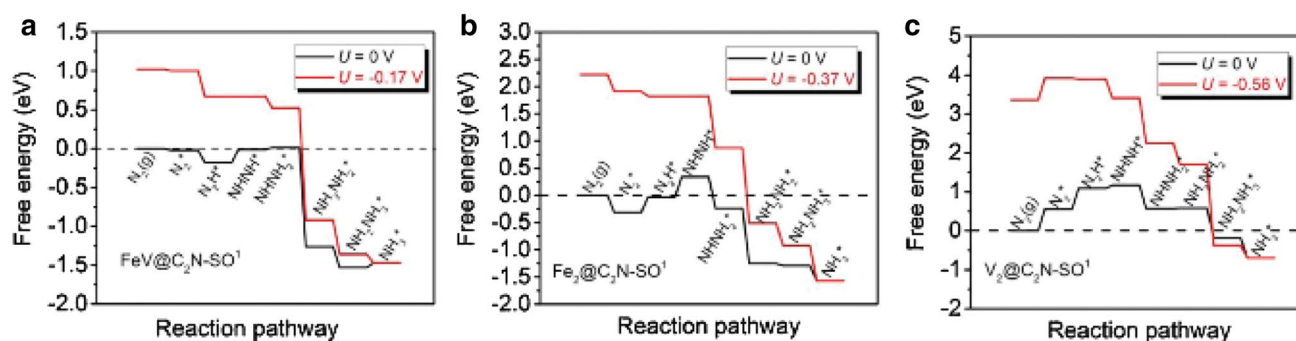
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HIGHLIGHTS

- A new heteronuclear bi-atom electrocatalyst has been proposed by Ma and his co-workers.
- The FeV@C₂N bi-atom electrocatalyst achieved excellent electrochemical NRR performance.
- The FeV@C₂N bi-atom electrocatalyst could effectively suppress the side and competing HER reaction, and thus possess better electrochemical NRR selectivity.

ABSTRACT The electrochemical nitrogen reduction reaction (NRR) to directly produce NH₃ from N₂ and H₂O under ambient conditions has attracted significant attention due to its ecofriendliness. Nevertheless, the electrochemical NRR presents several practical challenges, including sluggish reaction and low selectivity. Here, bi-atom catalysts have been proposed to achieve excellent activity and high selectivity toward the electrochemical NRR by Ma and his co-workers. It could accelerate the kinetics of N₂-to-NH₃ electrochemical conversion and possess better electrochemical NRR selectivity. This work sheds light on the introduction of bi-atom catalysts to enhance the performance of the electrochemical NRR.



KEYWORDS Electrochemical nitrogen reduction reaction; Bi-atom catalysts; Excellent activity; High selectivity

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The electrochemical nitrogen reduction reaction (NRR) to directly produce NH_3 from N_2 and H_2O under ambient conditions has attracted significant attention due to its eco-friendliness compared with the traditional Haber–Bosch process [1, 2]. Nevertheless, the electrochemical NRR presents several practical challenges, including sluggish reaction and low selectivity [3, 4]. The slow kinetics is caused by the extremely strong $\text{N}\equiv\text{N}$ triple bond (941 kJ mol^{-1}) and the great energy gap between highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO) of the N_2 molecule [5]. The hydrogen evolution reaction (HER) is the main side reaction responsible for the low selectivity, which shares a very close potential window with the NRR in both alkaline and acidic electrolytes [6, 7]. Fortunately, the electrochemical NRR depends heavily on its electrocatalysts [8–10]. Therefore, advanced rational design of the electrochemical NRR electrocatalysts to achieve outstanding performance and high selectivity is urgently required [11–13]. Various NRR electrocatalysts, including metal-free catalysts, single-atom catalysts, metal nanomaterials, nitrides/oxides/sulfides/carbides, etc., have been reported with the aim of high NH_3 yield since 2016 [14]. Nevertheless, a promising candidate, a heteronuclear bi-atom electrocatalyst, has been little studied for the electrochemical NRR.

Recently, Ma and co-workers [15] designed a new heteronuclear bi-atom electrocatalyst, Fe, V co-doped C_2N ($\text{FeV@C}_2\text{N}$), to accelerate the kinetics of the NRR and suppress the hydrogen evolution reaction (HER), which occurs as a side reaction. This $\text{FeV@C}_2\text{N}$ electrocatalyst achieved excellent electrochemical NRR performance. The nitrogenated holey structures in C_2N could anchor these Fe and V atoms; additionally, the unoccupied/occupied d orbitals of Fe and V atoms may accept/donate electrons from/to N_2 (Fig. 1a). Therefore, Fe and V atoms could be stable on the C_2N matrix and serve as active sites to electrocatalytically transform N_2 into NH_3 . The $\text{FeV@C}_2\text{N}$ could weaken the $\text{N}\equiv\text{N}$ triple bond and increase the Bader charge difference of two chemisorbed N atoms, as shown in Fig. 1b–d. More importantly, the $\text{FeV@C}_2\text{N}$ possesses the greatest ability to activate N_2 compared to $\text{Fe}_2\text@C}_2\text{N}$ and $\text{V}_2\text@C}_2\text{N}$.

Furthermore, Ma and co-workers [15] proposed the mechanism of N_2 reduction and free energy diagrams on side-on configurations of $\text{FeV@C}_2\text{N}$, $\text{Fe}_2\text@C}_2\text{N}$, and $\text{V}_2\text@C}_2\text{N}$, and

they believe that $\text{FeV@C}_2\text{N}$ is the most promising electrocatalyst for the NRR compared with the other two. There are only two reaction steps from N_2H^* to NHNH^* and NHNH^* to NHNH_2^* , which are endothermic for $\text{FeV@C}_2\text{N}$; thus, these two steps are the potential-determining step (PDS) with free energy of 0.17 eV, as shown in Fig. 2a. The $\text{Fe}_2\text@C}_2\text{N}$ shares the same PDS but with a higher free energy of 0.37 eV (Fig. 2b). In $\text{V}_2\text@C}_2\text{N}$, the PDS is the formation of N_2^* , with the greatest free energy of 0.56 eV (Fig. 2c). The HER, as a competing and side reaction, was also investigated for these three samples. The calculated results indicated that $\text{FeV@C}_2\text{N}$ and $\text{Fe}_2\text@C}_2\text{N}$ have better NRR selectivity. $\text{V}_2\text@C}_2\text{N}$, because of the competing HER, is not a good candidate for the electrochemical NRR.

In sum, Ma and co-workers proposed $\text{FeV@C}_2\text{N}$ as an outstanding heteronuclear bi-atom electrocatalyst for the electrochemical NRR, with high activity and better selectivity. It could enhance the kinetics of N_2 -to- NH_3 electrochemical conversion with a low potential PDS of -0.17 V. Moreover, this $\text{FeV@C}_2\text{N}$ electrocatalyst could effectively suppress the side and competing HER reaction, and thus possess better electrochemical NRR selectivity. This work sheds light on the introduction of heteronuclear bi-atom electrocatalysts to enhance the performance of the electrochemical NRR and opens a new way to understand the electrochemical NRR mechanism.

In the future, two possible prospects could be effective approaches to optimize the electrocatalysts with the aim of improved NRR activity and selectivity, and reveal the mechanisms of the electrochemical NRR as well. Firstly, theoretical calculations could be employed to predict potential NRR electrocatalysts and provide various types of optimization guidance to the experiments. For example, high-throughput computing can identify the poisoning and decomposition of electrocatalysts under electrochemical conditions, including the pH and the electrolyte effect, which can provide a deeper insight into the mechanism under real operation conditions. On the other hand, advanced characterization, including *in-situ* and *operando* atomic-resolution transmission electron microscopy and X-ray absorption spectroscopy, can be developed to identify the real active sites and

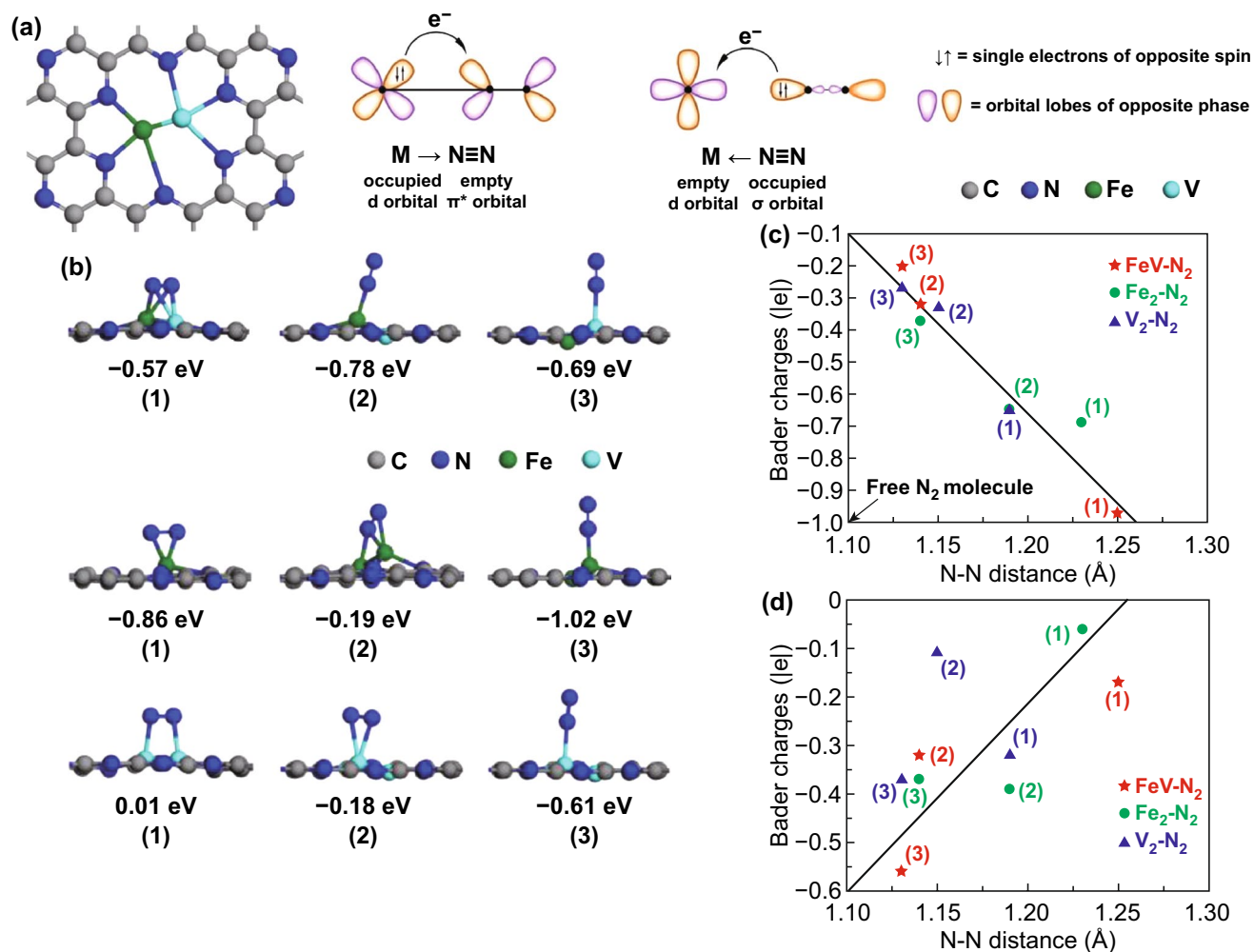


Fig. 1 **a** Optimized structures of FeV anchored on C₂N substrate and simplified schematic diagram of the bonding between the transition metal and N₂. **b** Optimized structures and corresponding adsorption energies of N₂ adsorption on FeV@C₂N, Fe₂@C₂N, and V₂@C₂N. **c** Relationship between Bader charges of adsorbed N₂ and N–N bond lengths. **d** Relationship between Bader charge difference of two adsorbed N atoms and N–N bond lengths [15]. Copyright 2020 Elsevier

composite evolution of the electrical double layer. With the significant efforts that have been made in the past few years, the electrochemical NRR appears promising to replace the traditional Haber–Bosch process to produce NH₃. Nevertheless, a reproducible and excellent electrochemical NRR catalyst is still expected to be proposed as a standard electrocatalyst, due to the doubt that has arisen on the actual

NRR performance. A benchmarking protocol to accurately quantify the electrochemical NRR activity and selectivity should be established. We believe that, with much effort, the fundamental issues and technological drawbacks will be addressed in the not-too-distant future, and the electrochemical NRR can play an important role in NH₃ yield.

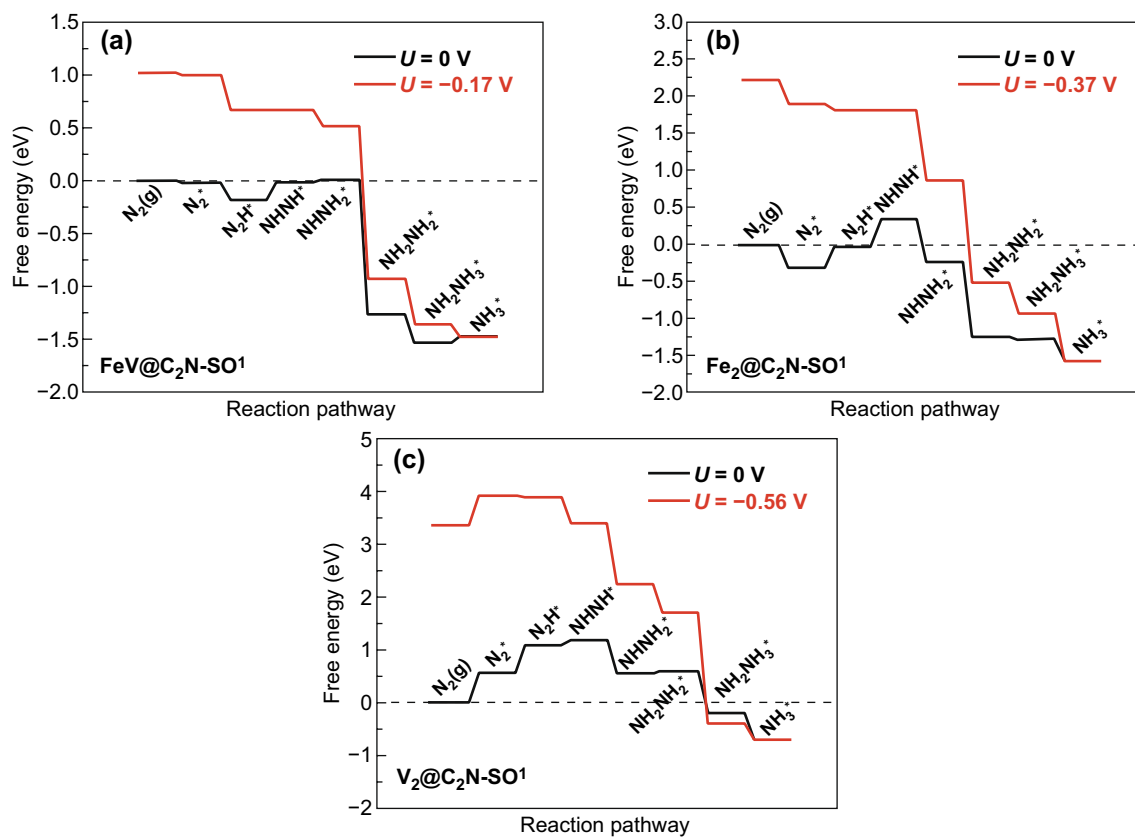


Fig. 2 a–c Free energy diagrams of the three electrocatalysts [15]. Copyright 2020 Elsevier

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